

## TABELA DE POTENCIAIS PADRÃO DE REDUÇÃO EM SOLUÇÃO AQUOSA A 25 °C

Soluções ácidas	$\epsilon^{\circ}$ (V)
$\text{H}_4\text{XeO}_6(\text{aq}) + 2 \text{H}^+(\text{aq}) + 2 \text{e}^- \rightleftharpoons \text{XeO}_3(\text{s}) + 3 \text{H}_2\text{O}$	+3,0
$\text{F}_2(\text{g}) + 2 \text{e}^- \rightleftharpoons 2 \text{F}^-(\text{aq})$	+2,87
$\text{O}_3(\text{g}) + 2 \text{H}^+(\text{aq}) + 2 \text{e}^- \rightleftharpoons \text{O}_2(\text{g}) + \text{H}_2\text{O}$	+2,07
$2\text{S}_2\text{O}_8^{2-}(\text{aq}) + 2 \text{e}^- \rightleftharpoons 2 \text{SO}_4^{2-}(\text{aq})$	+2,01
$\text{Co}^{3+}(\text{aq}) + \text{e}^- \rightleftharpoons \text{Co}^{2+}(\text{aq})$	+1,82
$\text{Pb}^{4+}(\text{aq}) + 2 \text{e}^- \rightleftharpoons \text{Pb}^{2+}(\text{aq})$	+1,8
$\text{H}_2\text{O}_2(\text{aq}) + 2 \text{H}^+(\text{aq}) + 2 \text{e}^- \rightleftharpoons 2 \text{H}_2\text{O}$	+1,776
$\text{NiO}_2(\text{s}) + 4 \text{H}^+(\text{aq}) + 2 \text{e}^- \rightleftharpoons \text{Ni}^{2+}(\text{aq}) + 2 \text{H}_2\text{O}$	+1,7
$\text{Pb}^{4+}(\text{aq}) + 2 \text{e}^- \rightleftharpoons \text{Pb}^{2+}(\text{aq})$	+1,70
$\text{MnO}_4^-(\text{aq}) + 4 \text{H}^+(\text{aq}) + 3 \text{e}^- \rightleftharpoons \text{MnO}_2(\text{s}) + 2 \text{H}_2\text{O}$	+1,679
$\text{PbO}_2(\text{s}) + \text{SO}_4^{2-}(\text{aq}) + 4 \text{H}^+(\text{aq}) + 2 \text{e}^- \rightleftharpoons \text{PbSO}_4(\text{s}) + 2 \text{H}_2\text{O}$	+1,685
$\text{Au}^+(\text{aq}) + \text{e}^- \rightleftharpoons \text{Au}(\text{s})$	+1,68
$2 \text{HClO}(\text{aq}) + 2 \text{H}^+(\text{aq}) + 2 \text{e}^- \rightleftharpoons \text{Cl}_2(\text{g}) + 2 \text{H}_2\text{O}$	+1,63
$\text{Ce}^{4+}(\text{aq}) + \text{e}^- \rightleftharpoons \text{Ce}^{3+}(\text{aq})$	+1,61
$\text{NaBiO}_3(\text{s}) + 6 \text{H}^+(\text{aq}) + 2 \text{e}^- \rightleftharpoons \text{Bi}^{3+}(\text{aq}) + \text{Na}^+(\text{aq}) + 3 \text{H}_2\text{O}$	+~1,6
$\text{MnO}_4^-(\text{aq}) + 8 \text{H}^+(\text{aq}) + 5 \text{e}^- \rightleftharpoons \text{Mn}^{2+}(\text{aq}) + 4 \text{H}_2\text{O}$	+1,51
$\text{Au}^{3+}(\text{aq}) + 3 \text{e}^- \rightleftharpoons \text{Au}(\text{s})$	+1,50
$\text{ClO}_3^-(\text{aq}) + 6 \text{H}^+(\text{aq}) + 5 \text{e}^- \rightleftharpoons \frac{1}{2} \text{Cl}_2(\text{g}) + 3 \text{H}_2\text{O}$	+1,47
$\text{PbO}_2(\text{s}) + 4 \text{H}^+(\text{aq}) + 2 \text{e}^- \rightleftharpoons \text{Pb}^{2+}(\text{aq}) + 2 \text{H}_2\text{O}$	+1,45
$\text{BrO}_3^- + 6 \text{H}^+(\text{aq}) + 6 \text{e}^- \rightleftharpoons \text{Br}^-(\text{aq}) + 3 \text{H}_2\text{O}$	+1,44
$\text{Cl}_2(\text{g}) + 2 \text{e}^- \rightleftharpoons 2 \text{Cl}^-(\text{aq})$	+1,358
$\text{Cr}_2\text{O}_7^{2-} + 14 \text{H}^+(\text{aq}) + 6 \text{e}^- \rightleftharpoons 2 \text{Cr}^{3+}(\text{aq}) + 7 \text{H}_2\text{O}$	+1,33
$\text{N}_2\text{H}_5^+(\text{aq}) + 3 \text{H}^+(\text{aq}) + 2 \text{e}^- \rightleftharpoons 2 \text{NH}_4^+(\text{aq})$	+1,24
$\text{MnO}_2(\text{s}) + 4 \text{H}^+(\text{aq}) + 2 \text{e}^- \rightleftharpoons \text{Mn}^{2+}(\text{aq}) + 2 \text{H}_2\text{O}$	+1,23
$\text{O}_2(\text{g}) + 4 \text{H}^+(\text{aq}) + 4 \text{e}^- \rightleftharpoons 2 \text{H}_2\text{O}$	+1,229
$\text{Pt}^{2+}(\text{aq}) + 2 \text{e}^- \rightleftharpoons \text{Pt}(\text{s})$	+1,2
$\text{IO}_3^-(\text{aq}) + 6 \text{H}^+(\text{aq}) + 5 \text{e}^- \rightleftharpoons \frac{1}{2} \text{I}_2(\text{aq}) + 3 \text{H}_2\text{O}$	+1,195
$\text{ClO}_4^-(\text{aq}) + 2 \text{H}^+(\text{aq}) + 2 \text{e}^- \rightleftharpoons \text{ClO}_3^-(\text{aq}) + \text{H}_2\text{O}$	+1,19
$\text{Br}_2(\text{l}) + 2 \text{e}^- \rightleftharpoons 2 \text{Br}^-(\text{aq})$	+1,066
$\text{AuCl}_4^- + 3 \text{e}^- \rightleftharpoons \text{Au}(\text{s}) + 4 \text{Cl}^-(\text{aq})$	+1
$\text{Pd}^{2+}(\text{aq}) + 2 \text{e}^- \rightleftharpoons \text{Pd}(\text{s})$	+0,987
$\text{NO}_3^-(\text{aq}) + 4 \text{H}^+(\text{aq}) + 3 \text{e}^- \rightleftharpoons \text{NO}(\text{g}) + 2 \text{H}_2\text{O}$	+0,96
$\text{NO}_3^-(\text{aq}) + 3 \text{H}^+(\text{aq}) + 2 \text{e}^- \rightleftharpoons \text{HNO}_2(\text{aq}) + \text{H}_2\text{O}$	+0,94
$2 \text{Hg}^{2+}(\text{aq}) + 2 \text{e}^- \rightleftharpoons \text{Hg}_2^{2+}(\text{aq})$	+0,92
$\text{Hg}^{2+}(\text{aq}) + 2 \text{e}^- \rightleftharpoons \text{Hg}(\text{l})$	+0,855

$2\text{NO}_3^-(\text{aq}) + 4\text{H}^+(\text{aq}) + 2\text{e}^- \rightleftharpoons 2\text{NO}_2(\text{g}) + 2\text{H}_2\text{O}$	+0,80
$\text{Ag}^+(\text{aq}) + \text{e}^- \rightleftharpoons \text{Ag}(\text{s})$	+0,7994
$\text{Hg}_2^{2+}(\text{aq}) + 2\text{e}^- \rightleftharpoons 2\text{Hg}(\text{l})$	+0,789
$\text{Fe}^{3+}(\text{aq}) + \text{e}^- \rightleftharpoons \text{Fe}^{2+}(\text{aq})$	+0,771
$\text{SbCl}_6^-(\text{aq}) + 2\text{e}^- \rightleftharpoons \text{SbCl}_4^-(\text{aq}) + 2\text{Cl}^-(\text{aq})$	+0,75
$[\text{PtCl}_4]^{2-}(\text{aq}) + 2\text{e}^- \rightleftharpoons \text{Pt}(\text{s}) + 4\text{Cl}^-(\text{aq})$	+0,73
$\text{O}_2(\text{g}) + 2\text{H}^+(\text{aq}) + 2\text{e}^- \rightleftharpoons \text{H}_2\text{O}_2(\text{aq})$	+0,682
$[\text{PtCl}_6]^{2-}(\text{aq}) + 2\text{e}^- \rightleftharpoons [\text{PtCl}_4]^{2-}(\text{aq}) + 2\text{Cl}^-(\text{aq})$	+0,68
$\text{H}_3\text{AsO}_4(\text{aq}) + 2\text{H}^+(\text{aq}) + 2\text{e}^- \rightleftharpoons \text{H}_3\text{AsO}_3(\text{aq}) + \text{H}_2\text{O}$	+0,58
$\text{H}_3\text{AsO}_4(\text{aq}) + 2\text{H}^+(\text{aq}) + 2\text{e}^- \rightleftharpoons \text{H}_3\text{AsO}_2(\text{aq}) + 2\text{H}_2\text{O}$	+0,560
$\text{MnO}_4^-(\text{aq}) + \text{e}^- \rightleftharpoons \text{MnO}_4^{2-}(\text{aq})$	+0,558
$\text{I}_3^-(\text{aq}) + 2\text{e}^- \rightleftharpoons 3\text{I}^-(\text{aq})$	+0,536
$\text{I}_2(\text{s}) + 2\text{e}^- \rightleftharpoons 2\text{I}^-(\text{aq})$	+0,535
$\text{TeO}_2(\text{s}) + 4\text{H}^+(\text{aq}) + 4\text{e}^- \rightleftharpoons \text{Te}(\text{s}) + 2\text{H}_2\text{O}$	+0,529
$\text{Cu}^+(\text{aq}) + \text{e}^- \rightleftharpoons \text{Cu}(\text{s})$	+0,521
$[\text{RhCl}_6]^{3-}(\text{aq}) + 3\text{e}^- \rightleftharpoons \text{Rh}(\text{s}) + 6\text{Cl}^-(\text{aq})$	+0,44
$\text{Fe}(\text{CN})_6^{3-}(\text{aq}) + \text{e}^- \rightleftharpoons \text{Fe}(\text{CN})_6^{4-}(\text{aq})$	+0,36
$\text{Cu}^{2+}(\text{aq}) + 2\text{e}^- \rightleftharpoons \text{Cu}(\text{s})$	+0,337
$\text{HgCl}_2(\text{s}) + 2\text{e}^- \rightleftharpoons 2\text{Hg}(\text{l}) + 2\text{Cl}^-(\text{aq})$	+0,27
$\text{AgCl}(\text{s}) + \text{e}^- \rightleftharpoons \text{Ag}(\text{s}) + \text{Cl}^-(\text{aq})$	+0,222
$\text{SO}_4^{2-}(\text{aq}) + 4\text{H}^+(\text{aq}) + 2\text{e}^- \rightleftharpoons \text{SO}_2(\text{g}) + 2\text{H}_2\text{O}$	+0,2
$\text{SO}_4^{2-}(\text{aq}) + 4\text{H}^+(\text{aq}) + 2\text{e}^- \rightleftharpoons \text{H}_2\text{SO}_3(\text{g}) + \text{H}_2\text{O}$	+0,17
$\text{Cu}^{2+}(\text{aq}) + \text{e}^- \rightleftharpoons \text{Cu}^+(\text{aq})$	+0,153
$\text{Sn}^{4+}(\text{aq}) + 2\text{e}^- \rightleftharpoons \text{Sn}^{2+}(\text{aq})$	+0,15
$\text{S}(\text{s}) + 2\text{H}^+(\text{aq}) + 2\text{e}^- \rightleftharpoons \text{H}_2\text{S}(\text{aq})$	+0,14
$\text{AgBr}(\text{s}) + \text{e}^- \rightleftharpoons \text{Ag}(\text{s}) + \text{Br}^-(\text{aq})$	+0,0713
$2\text{H}^+(\text{aq}) + 2\text{e}^- \rightleftharpoons \text{H}_2(\text{g})$ (eletrodo de referência; definição)	0
$2\text{D}^+(\text{aq}) + 2\text{e}^- \rightleftharpoons \text{D}_2(\text{g})$	-0,0034
$\text{Fe}^{3+}(\text{aq}) + 3\text{e}^- \rightleftharpoons \text{Fe}(\text{s})$	-0,037
$\text{N}_2\text{O}(\text{g}) + 6\text{H}^+(\text{aq}) + \text{H}_2\text{O} + 4\text{e}^- \rightleftharpoons 2\text{NH}_3\text{OH}^+(\text{aq})$	-0,05
$\text{Pb}^{2+}(\text{aq}) + 2\text{e}^- \rightleftharpoons \text{Pb}(\text{s})$	-0,126
$\text{Sn}^{2+}(\text{aq}) + 2\text{e}^- \rightleftharpoons \text{Sn}(\text{s})$	-0,14
$\text{AgI}(\text{s}) + \text{e}^- \rightleftharpoons \text{Ag}(\text{s}) + \text{I}^-(\text{aq})$	-0,15
$[\text{SnF}_6]^{2-}(\text{aq}) + 4\text{e}^- \rightleftharpoons \text{Sn}(\text{s}) + 6\text{F}^-(\text{aq})$	-0,25
$\text{Ni}^{2+}(\text{aq}) + 2\text{e}^- \rightleftharpoons \text{Ni}(\text{s})$	-0,25
$\text{Co}^{2+}(\text{aq}) + 2\text{e}^- \rightleftharpoons \text{Co}(\text{s})$	-0,28
$\text{Tl}^+(\text{aq}) + \text{e}^- \rightleftharpoons \text{Tl}(\text{s})$	-0,34
$\text{PbSO}_4(\text{s}) + 2\text{e}^- \rightleftharpoons \text{Pb}(\text{s}) + \text{SO}_4^{2-}(\text{aq})$	-0,356

$\text{Se(s)} + 2 \text{H}^+(\text{aq}) + 2 \text{e}^- \rightleftharpoons \text{H}_2\text{Se(aq)}$	-0,4
$\text{Cd}^{2+}(\text{aq}) + 2 \text{e}^- \rightleftharpoons \text{Cd(s)}$	-0,403
$\text{Cr}^{3+}(\text{aq}) + \text{e}^- \rightleftharpoons \text{Cr}^{2+}(\text{aq})$	-0,41
$\text{Fe}^{2+}(\text{aq}) + 2 \text{e}^- \rightleftharpoons \text{Fe(s)}$	-0,44
$\text{S(s)} + 2 \text{e}^- \rightleftharpoons \text{S}^{2-}(\text{aq})$	-0,48
$2 \text{CO}_2(\text{g}) + 2 \text{H}^+(\text{aq}) + 2 \text{e}^- \rightleftharpoons (\text{COOH})_2(\text{aq})$	-0,49
$\text{Ga}^{3+}(\text{aq}) + 3 \text{e}^- \rightleftharpoons \text{Ga(s)}$	-0,53
$\text{HgS(s)} + 2 \text{H}^+(\text{aq}) + 2 \text{e}^- \rightleftharpoons \text{Hg(l)} + \text{H}_2\text{S(g)}$	-0,72
$\text{Cr}^{3+}(\text{aq}) + 3 \text{e}^- \rightleftharpoons \text{Cr(s)}$	-0,74
$\text{Zn}^{2+}(\text{aq}) + 2 \text{e}^- \rightleftharpoons \text{Zn(s)}$	-0,763
$2 \text{H}_2\text{O(l)} + 2 \text{e}^- \rightleftharpoons \text{H}_2(\text{g}) + 2 \text{OH}^-(\text{aq})$	-0,8277
$\text{Cr}^{2+}(\text{aq}) + 2 \text{e}^- \rightleftharpoons \text{Cr(s)}$	-0,91
$\text{Mn}^{2+}(\text{aq}) + 2 \text{e}^- \rightleftharpoons \text{Mn(s)}$	-1,18
$\text{V}^{2+}(\text{aq}) + 2 \text{e}^- \rightleftharpoons \text{V(s)}$	-1,19
$\text{SiF}_6^{2-}(\text{aq}) + 4 \text{e}^- \rightleftharpoons \text{Si(s)} + 6 \text{F}^-(\text{aq})$	-1,24
$\text{Zr}^{4+}(\text{aq}) + 4 \text{e}^- \rightleftharpoons \text{Zr(s)}$	-1,53
$\text{Al}^{3+}(\text{aq}) + 3 \text{e}^- \rightleftharpoons \text{Al(s)}$	-1,66
$\text{AlF}_6^{3-}(\text{aq}) + 3 \text{e}^- \rightleftharpoons \text{Al(s)} + 6 \text{F}^-(\text{aq})$	-2,07
$\text{H}_2(\text{g}) + 2 \text{e}^- \rightleftharpoons 2 \text{H}^-(\text{aq})$	-2,25
$\text{Mg}^{2+}(\text{aq}) + 2 \text{e}^- \rightleftharpoons \text{Mg(s)}$	-2,37
$\text{La}^{3+}(\text{aq}) + 3 \text{e}^- \rightleftharpoons \text{La(s)}$	-2,52
$\text{Na}^+(\text{aq}) + \text{e}^- \rightleftharpoons \text{Na(s)}$	-2,714
$\text{Ca}^{2+}(\text{aq}) + 2 \text{e}^- \rightleftharpoons \text{Ca(s)}$	-2,87
$\text{Sr}^{2+}(\text{aq}) + 2 \text{e}^- \rightleftharpoons \text{Sr(s)}$	-2,89
$\text{Ba}^{2+}(\text{aq}) + 2 \text{e}^- \rightleftharpoons \text{Ba(s)}$	-2,9
$\text{Rb}^+(\text{aq}) + \text{e}^- \rightleftharpoons \text{Rb(s)}$	-2,925
$\text{K}^+(\text{aq}) + \text{e}^- \rightleftharpoons \text{K(s)}$	-2,925
$\text{Li}^+(\text{aq}) + \text{e}^- \rightleftharpoons \text{Li(s)}$	-3,045

### Soluções básicas

	$\epsilon^{\circ}(\text{V})$
$\text{HXeO}_4^-(\text{aq}) + 3 \text{H}_2\text{O} + 6 \text{e}^- \rightleftharpoons \text{Xe(g)} + 7 \text{OH}^-(\text{aq})$	+0,9
$\text{ClO}^-(\text{aq}) + \text{H}_2\text{O} + 2 \text{e}^- \rightleftharpoons \text{Cl}^-(\text{aq}) + 2 \text{OH}^-(\text{aq})$	+0,89
$\text{OOH}^-(\text{aq}) + \text{H}_2\text{O} + 2 \text{e}^- \rightleftharpoons 3 \text{OH}^-(\text{aq})$	+0,88
$2 \text{NH}_2\text{OH(aq)} + 2 \text{e}^- \rightleftharpoons \text{N}_2\text{H}_4(\text{aq}) + 2 \text{OH}^-(\text{aq})$	+0,74
$\text{ClO}_3^-(\text{aq}) + 3 \text{H}_2\text{O} + 6 \text{e}^- \rightleftharpoons \text{Cl}^-(\text{aq}) + 6 \text{OH}^-(\text{aq})$	+0,62
$\text{MnO}_4^-(\text{aq}) + 2 \text{H}_2\text{O} + 3 \text{e}^- \rightleftharpoons \text{MnO}_2(\text{s}) + 4 \text{OH}^-(\text{aq})$	+0,588
$\text{MnO}_4^-(\text{aq}) + \text{e}^- \rightleftharpoons \text{MnO}_4^{2-}(\text{aq})$	+0,564
$\text{NiO}_2(\text{s}) + 2 \text{H}_2\text{O} + 2 \text{e}^- \rightleftharpoons \text{Ni(OH)}_2(\text{s}) + 2 \text{OH}^-(\text{aq})$	+0,49

$\text{Ag}_2\text{CrO}_4(\text{s}) + 2 \text{e}^- \rightleftharpoons 2 \text{Ag}(\text{s}) + \text{CrO}_4^{2-}(\text{aq})$	+0,446
$\text{O}_2(\text{g}) + 2 \text{H}_2\text{O} + 4 \text{e}^- \rightleftharpoons 4 \text{OH}^-(\text{aq})$	+0,4
$\text{ClO}_4^-(\text{aq}) + \text{H}_2\text{O} + 2 \text{e}^- \rightleftharpoons \text{ClO}_3^-(\text{aq}) + 2 \text{OH}^-(\text{aq})$	+0,36
$\text{Ag}_2\text{O}(\text{s}) + \text{H}_2\text{O} + 2 \text{e}^- \rightleftharpoons 2 \text{Ag}(\text{s}) + 2 \text{OH}^-(\text{aq})$	+0,34
$\text{ClO}_3^-(\text{aq}) + \text{H}_2\text{O} + 2 \text{e}^- \rightleftharpoons \text{ClO}_2^-(\text{aq}) + 2 \text{OH}^-(\text{aq})$	+0,33
$\text{IO}_3^-(\text{aq}) + 3 \text{H}_2\text{O} + 6 \text{e}^- \rightleftharpoons \text{I}^-(\text{aq}) + 6 \text{OH}^-(\text{aq})$	+0,26
$\text{Co}(\text{OH})_3(\text{aq}) + \text{e}^- \rightleftharpoons \text{Co}(\text{OH})_2(\text{aq}) + \text{OH}^-(\text{aq})$	+0,17
$2 \text{NO}_2^-(\text{aq}) + 3 \text{H}_2\text{O} + 4 \text{e}^- \rightleftharpoons \text{N}_2\text{O}(\text{g}) + 6 \text{OH}^-(\text{aq})$	+0,15
$\text{N}_2\text{H}_4(\text{aq}) + 2 \text{H}_2\text{O} + 2 \text{e}^- \rightleftharpoons 2 \text{NH}_3(\text{aq}) + 2 \text{OH}^-(\text{aq})$	+0,1
$[\text{Co}(\text{NH}_3)_6]^{3+}(\text{aq}) + \text{e}^- \rightleftharpoons [\text{Co}(\text{NH}_3)_6]^{2+}(\text{aq})$	+0,1
$\text{HgO}(\text{s}) + \text{H}_2\text{O} + 2 \text{e}^- \rightleftharpoons \text{Hg}(\text{l}) + 2 \text{OH}^-(\text{aq})$	+0,0984
$\text{O}_2(\text{g}) + \text{H}_2\text{O} + 2 \text{e}^- \rightleftharpoons \text{OOH}^-(\text{aq}) + \text{OH}^-(\text{aq})$	+0,076
$\text{NO}_3^-(\text{aq}) + \text{H}_2\text{O} + 2 \text{e}^- \rightleftharpoons \text{NO}_2^-(\text{aq}) + 2 \text{OH}^-(\text{aq})$	+0,01
$\text{MnO}_2(\text{s}) + 2 \text{H}_2\text{O} + 2 \text{e}^- \rightleftharpoons \text{Mn}(\text{OH})_2(\text{s}) + 2 \text{OH}^-(\text{aq})$	-0,05
$\text{CrO}_4^{2-}(\text{aq}) + 4 \text{H}_2\text{O} + 3 \text{e}^- \rightleftharpoons \text{Cr}(\text{OH})_3(\text{s}) + 5 \text{OH}^-(\text{aq})$	-0,12
$\text{Cu}(\text{OH})_2(\text{s}) + 2 \text{e}^- \rightleftharpoons \text{Cu}(\text{s}) + 2 \text{OH}^-(\text{aq})$	-0,36
$\text{HPbO}_2^-(\text{aq}) + \text{H}_2\text{O} + 2 \text{e}^- \rightleftharpoons \text{Pb}(\text{s}) + 3 \text{OH}^-(\text{aq})$	-0,54
$\text{Fe}(\text{OH})_3(\text{s}) + \text{e}^- \rightleftharpoons \text{Fe}(\text{OH})_2(\text{s}) + \text{OH}^-(\text{aq})$	-0,56
$2 \text{H}_2\text{O} + 2 \text{e}^- \rightleftharpoons \text{H}_2(\text{g}) + 2 \text{OH}^-(\text{aq})$	-0,8277
$2 \text{NO}_3^-(\text{aq}) + 2 \text{H}_2\text{O} + 2 \text{e}^- \rightleftharpoons \text{N}_2\text{O}_4(\text{g}) + 4 \text{OH}^-(\text{aq})$	-0,85
$\text{Fe}(\text{OH})_2(\text{s}) + 2 \text{e}^- \rightleftharpoons \text{Fe}(\text{s}) + 2 \text{OH}^-(\text{aq})$	-0,877
$\text{HSnO}_2^-(\text{aq}) + \text{H}_2\text{O} + 2 \text{e}^- \rightleftharpoons \text{Sn}(\text{s}) + 3 \text{OH}^-(\text{aq})$	-0,91
$\text{Sn}(\text{OH})_6^{2-}(\text{aq}) + 2 \text{e}^- \rightleftharpoons \text{HSnO}_2^-(\text{aq}) + 2 \text{H}_2\text{O} + 3 \text{OH}^-(\text{aq})$	-0,93
$\text{SO}_4^{2-}(\text{aq}) + \text{H}_2\text{O} + 2 \text{e}^- \rightleftharpoons \text{SO}_3^{2-}(\text{aq}) + 2 \text{OH}^-(\text{aq})$	-0,93
$\text{N}_2(\text{g}) + 4 \text{H}_2\text{O} + 4 \text{e}^- \rightleftharpoons \text{N}_2\text{H}_4(\text{aq}) + 4 \text{OH}^-(\text{aq})$	-1,15
$\text{ZnO}_2^{2-}(\text{aq}) + 2 \text{H}_2\text{O} + 2 \text{e}^- \rightleftharpoons \text{Zn}(\text{s}) + 4 \text{OH}^-(\text{aq})$	-1,21
$[\text{Zn}(\text{OH})_4]^{2-}(\text{aq}) + 2 \text{e}^- \rightleftharpoons \text{Zn}(\text{s}) + 4 \text{OH}^-(\text{aq})$	-1,22
$\text{Zn}(\text{OH})_2(\text{s}) + 2 \text{e}^- \rightleftharpoons \text{Zn}(\text{s}) + 2 \text{OH}^-(\text{aq})$	-1,245
$[\text{Zn}(\text{CN})_4]^{2-}(\text{aq}) + 2 \text{e}^- \rightleftharpoons \text{Zn}(\text{s}) + 4 \text{CN}^-(\text{aq})$	-1,26
$\text{CrO}_2^-(\text{aq}) + 2 \text{H}_2\text{O} + 3 \text{e}^- \rightleftharpoons \text{Cr}(\text{s}) + 4 \text{OH}^-(\text{aq})$	-1,27
$\text{Cr}(\text{OH})_3(\text{s}) + 3 \text{e}^- \rightleftharpoons \text{Cr}(\text{s}) + 3 \text{OH}^-(\text{aq})$	-1,3
$\text{SiO}_3^{2-}(\text{aq}) + 3 \text{H}_2\text{O} + 4 \text{e}^- \rightleftharpoons \text{Si}(\text{s}) + 6 \text{OH}^-(\text{aq})$	-1,7
$\text{H}_2\text{AlO}_3^-(\text{aq}) + 2 \text{H}_2\text{O} + 3 \text{e}^- \rightleftharpoons \text{Al}(\text{s}) + 4 \text{OH}^-(\text{aq})$	-2,33